Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**Student Activity- How Do Buffers Work?**

**Learning Objectives: SAP-10.B** Explain the relationship between the ability of a buffer to stabilize pH and the reactions that occur when an acid or a base is added to a buffered solution.

 **SAP-10.C** Identify the pH of a buffer solution based on the identity and concentrations of the conjugate acid-base pair used to create the buffer.

 **SAP-10.D** Explain the relationship between the buffer capacity of a solution and the relative concentrations of the conjugate acid and conjugate base components of the solution.

**Science Practice 6.D** Provide reasoning to justify a claim using chemical principles or laws, or using mathematical justification.

 **5.F** Calculate, estimate, or predict an unknown quantity from known quantities by selecting and following a logical computational pathway and attending to precision (e.g., performing dimensional analysis and attending to significant figures).

 **2.C** Identify experimental procedures that are aligned to a scientific question.

**QUESTION:** How do buffers wok?

**PROCEDURE:**

**PART I - TESTING THE AFFECT OF ACID AND BASE ON DISTILLED WATER**

1. Measure out exactly 25 mL of fresh DW into a beaker. Add a stir bar and place the beaker on top of a stir plate. Turn on the stirrer.

2. Insert pH paper into beaker and use color chart to determine pH. Record the pH.

3. Add exactly 10 drops of 0.10 M HCl. Determine the pH using a new piece of pH paper.

4. Measure out another 25 mL of fresh DW into a beaker. Add a stir bar and place the beaker on top of a stir plate. Turn on the stirrer.

5. Insert pH paper into beaker and use color chart to determine pH. Record the pH.

6. Add exactly 10 drops of 0.10 M NaOH. Determine the pH using a new piece of pH paper.

**See Data Table on next page.**

**PART II - TESTING THE AFFECT OF ACID AND BASE ON A WEAK ACID**

1. Measure out exactly 25 mL of 0.10 M acetic acid into a beaker. Add a stir bar and place the beaker on top of a stir plate. Turn on the stirrer.

2. Insert pH paper into beaker and use color chart to determine pH. Record the pH.

3. Add exactly 10 drops of 0.10 M HCl. Determine the pH using a new piece of pH paper.

4. Measure out another 25 mL of 0.10 M acetic acid into a beaker. Add a stir bar and place the beaker on top of a stir plate. Turn on the stirrer.

5 Insert pH paper into beaker and use color chart to determine pH. Record the pH.

6. Add exactly 10 drops of 0.10 M NaOH. Determine the pH using a new piece of pH paper.

**PART III - TESTING THE AFFECT OF ACID AND BASE ON A BUFFER**

1. Measure out exactly 25 mL of a buffer solution that contains 0.10 M acetic acid mixed with 0.10 M sodium acetate into a beaker. Add a stir bar and place the beaker on top of a stir plate. Turn on the stirrer.

2. Insert pH paper into beaker and use color chart to determine pH. Record the pH.

3. Add exactly 10 drops of 0.10 M HCl. Determine the pH using a new piece of pH paper.

4. Measure out another 25 mL of buffer solution into a beaker. Add a stir bar and place the beaker on top of a stir plate. Turn on the stirrer.

5. Insert pH paper into beaker and use color chart to determine pH. Record the pH.

6. Add exactly 10 drops of 0.10 M NaOH. Determine the pH using a new piece of pH paper.

**DATA: pH Data**

 **pH pH pH**

 **(no additions) (with 10 drops of HCl) (with 10 drops of NaOH)**

**Distilled Water \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_**

**Acetic Acid \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_**

**Buffer Solution \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_**

**QUESTIONS:**

1.Explain why pH changes when you add acid or base to water.

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**2.** Why did pH decrease when you added 0.10 M HCl to 0.10 M acetic acid?

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3. Write a net ionic equation showing the reaction between NaOH and HC2H3O2.

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4. Assume that 10 drops equals 1.0 mL. Calculate what the pH should have been when you added the 10 drops of 0.10 M NaOH to your 25 mL of Acetic Acid.

 (Acetic acid has a *Ka* of 1.8 x10-5).

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5. Write a net ionic equation showing the reaction between HCl and your

 buffer solution (HC2H3O2 and Na C2H3O2). Then calculate what the pH should have been when you added the 10 drops of 0.10 M HCl to your 25 mL of buffer.

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6. Write a net ionic equation showing the reaction between NaOH and your

 buffer solution (HC2H3O2 and Na C2H3O2). Then calculate what the pH should have been when you added the 10 drops of 0.10 M NaOH to your 25 mL of buffer.

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7. How close were your actual results to your theoretical results from questions 4, 5 and 6?

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8**.** Ammonia has a K*b* of 1.8 x 10-5. Design a buffer using 0.1 M solutions of ammonia and a

 0.1 M solution of another substance. Describe the other substance. What would be the pH of this buffer. Show your work.

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